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Stoichiometry MIXED  
REVIEW SHORT  
ANSWER Answer the following questions in the space provided. 1.

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Stoichiometry  
Answers Section 1

Given the following equation:  $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient  $x$  in this equation? 40.07 g/mol

b. What is the molar mass of  $C_3H_4$ ? 2 mol  $O_2$ :1 mol  $H_2O$

c. What is the mole ratio of  $O_2$  to  $H$

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#1 - #18, #31, & #38  
Answers . 38. To  
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## Stoichiometry Answers Section 1

magnesium is converted to  $MgO$ , I would use pure oxygen, not air, to carry out the reaction, because  $Mg$  could react with  $N_2$  in air to form  $Mg_3N_2$ . The pure oxygen should be used.

1 - 18, 31, & 38 Answers

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Modern Chemistry 77  
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CHAPTER 9 REVIEW

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## Stoichiometry SECTION 3 PROBLEMS Write the

answer on the line to the left. Show all your work in the space provided. 1. \_\_\_\_\_

The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2.

6.0 mol of  $N_2$  are mixed with 12.0 mol of H

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ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $C_3H_4(g) + x O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$  a. What is the value of the coefficient . x. in this equation? b.

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Answers Section 1

What is the molar mass of  $C_3H_4$ ? c. How many moles are in an 8.0 g sample of  $C_3H_4$ ? 2. a. What is meant by . ideal conditions

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answer on the line to

the left Show all your

work in the space

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## Stoichiometry Answers Section 1

### Answers

Stoichiometry b.

Theoretically, how many moles of  $\text{NH}_3$  will be produced?

PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2.

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Stoichiometry  
Answers Section 1  
6.0 mol of  $N_2$  are  
mixed with 12.0 mol of  
 $H_2$  according to the ...

## **Date. FCHAP] REV[EW.**

fewer steps are  
required to solve  
stoichiometry problems  
when. ... Chemistry

Chapter 9

Stoichiometry Test

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Chapter 6: Chemical

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following questions in  
the space provided. 1. 1

Given the following  
equation:  $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$   
4 a. What is  
the value of the  
coefficient x in this  
equation? 40.07 g/mol  
b.

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Stoichiometry &  
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**Stoichiometry &**  
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Answers Section 1

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Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry ... 290

Chapter 9 DO NOT

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Avogadro's number is:

Q. Using the pictured equation, how many grams of zinc chloride are produced from 7.89

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moles of zinc?

## Answers Section 1

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## Stoichiometry. 9-1 Introduction to Section 1

Stoichiometry.

Composition

Stoichiometry - deals  
with mass relationships  
of elements in

compounds Reaction

Stoichiometry -

Involves mass  
relationships between  
reactants and products  
in a chemical reaction.

I. Reaction

Stoichiometry

Problems A. Four

problem Types, One

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Stoichiometry  
Common Solution.

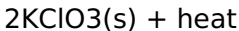
## Answers Section 1

### **Chapter 9 - Stoichiometry**

Ch. 9 Review:

Stoichiometry KEY

Page 1 1. The following equation represents a laboratory preparation for oxygen gas:



How many moles of  $\text{O}_2$  form as 3.0 mol of  $\text{KClO}_3$  are totally consumed?

$$3.0 \text{ mol KClO}_3 \times (3 \text{ moles O}_2) / (2 \text{ moles KClO}_3) =$$

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Stoichiometry  
4.5 moles O<sub>2</sub> 2.

## Answers Section 1

**Ch 9 Packet KEY |  
Stoichiometry | Mole  
(Unit) | Free 30 ...**

Chapter 9 -

Stoichiometry Chapter  
9 focuses on reaction  
stoichiometry: using a  
balanced chemical  
equation to calculate  
the number of grams,  
moles, or particles of  
reactants/products  
involved in a...

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