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In their excited state electrons are unstable and want to return to ground

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state. To do this, the electrons must lose the same quantum of energy they gained. They lose this energy in the form of LIGHT and/or heat. Once the energy is released, the electrons return to ground state

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andrewpham223. Key Concepts: Terms  
in this set (35) Dalton's atomic model.  
Atoms are indivisible and indestructible.

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All atoms of an element are identical.  
Atoms of different elements have  
differing sizes and masses.

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are the wavelength and frequency of light related? Wavelength. The . amplitude. of a wave is the wave's height from zero to the crest. The . wavelength, represented by  $\lambda$  (the Greek letter lambda), is the distance between the crests. It is often measured in meters or nanometers. Frequency

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Electrons move in circular orbits . around

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. the nucleus . at . fixed. energy. levels.  
Electrons are never between energy levels or energy shells. An electron must have . just the right amount of energy. to jump from one level to another. A . ...  
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valence electrons. Nitrogen has 3 electrons in 3 bonds and

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electrons in atoms. Electrons must have a certain minimum amount of energy called. In bohrs model of the atom electrons are in certain levels with the levels closest to the nucleus of energy than ...

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 $\text{s}^{-1}$  What is the energy of one quantum of  
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7s, 2p, 3p, 4p, 5p, 6p, 7p, 3d, CHAPTER  
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Both atoms have a filled s subshell outside their filled inner shells.

Aluminum (atomic number 13), with 13 electrons and the electron configuration

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[Ne]3s<sup>2</sup> 3p<sup>1</sup>, is analogous to its family member boron, [He]2s<sup>2</sup> 2p<sup>1</sup>.

## **2.2: Many-Electron Atoms and the Periodic Table ...**

Electrons in Atoms & Periodic Table  
Chapter 13 & 14 Assignment & Problem Set 7. An atom of an element has two electrons in the first energy level and

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five electrons in the second energy level.

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The electrons of the atoms are



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9. Electrons must have a certain minimum amount of energy called a quantum in order to move from one energy level to the next higher energy level. F 10. The electron probability clouds for atomic orbitals are spherical in shape. Chapter 5 Electrons in Atoms

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