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The big idea for all calorimetry processes is energy is conserved. Energy cannot be created or destroyed, but it can be exchanged. q lost + q gain = 0 or q released + q gain = 0. Students experience difficulty learning thermochemistry concepts. prove determine the heat exchanged at constant pressure, q = m c ΔT.

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Part 3: Calorimetry shows a basic diagram of a calorimeter, used to measure changes in temperature when combining objects. The Conclusions section contains four bullet points which must be addressed in their conclusion. At this point, I release students to spread out and work.